
EXAMPLE: You must be Psychro!

A moist-air mixture has a dry-bulb temperature of 85°F and a relative humidity of $\phi = 60\%$. The total pressure of the mix is 14.7 psia.

- (a) If the water vapor existed alone at T_{mix} and ψ_{mix} , what would its pressure be? I.e., determine the

- (b) For every lbm of dry air, how much water vapor is there? I.e., determine the

- (c) If you cooled this mix at constant pressure, at what temperature would the water start condensing? I.e., determine the

- (d) Determine the enthalpy of the mixture **per unit mass of dry air**. Is this the same as H_{mix}/m_{mix} ?